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$\Delta H = m \times C \times \Delta T$. Where M is the mass or volume of the solution being used to test the energy change, C is the Specific Heat Capacity of the solution, (for example the specific heat capacity of water is 4.1855 J g⁻¹ K⁻¹), and ΔT is the change in temperature.

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42. Given that heat neutralization of strong acid and strong base is -13.7 kcal. The heat produced when one mole of HCl is mixed with 0.5 mole of NaOH will be _____

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